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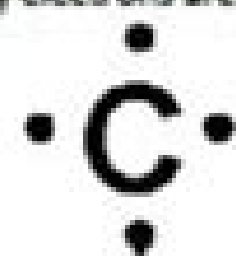
# Chemical Bonding Review

## Chemical Bonding Review

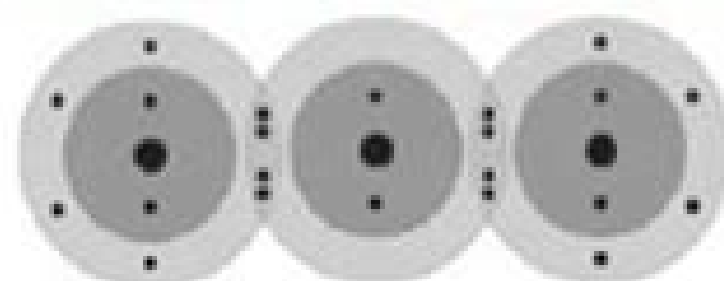
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Directions: Answer each of the following questions by circling the correct answer choice.

- Which two elements have the same number of valence electrons?
  - Boron and Aluminum
  - Carbon and Nitrogen
  - Argon and Helium
  - Silicon and Aluminum
- What type of bond forms when electrons are transferred?
  - Covalent
  - Ionic
  - Metallic
  - Nonmetallic
- What is the chemical symbol for a potassium ion?
  - $K^{+1}$
  - $K^{-1}$
  - $P^{+1}$
  - $P^{-1}$
- How many electrons are needed for a filled outermost energy level in the Lewis-dot structure below?



- 4
  - 0
  - 2
  - 1
- When atoms gain electrons during chemical bonding, they become
    - Positively charged ions
    - Neutral ions
    - Negatively charged ions
    - Chemically charged ions
  - Potassium chloride is an example of what type of bond?
    - Metallic bond
    - Colloid bond
    - Covalent bond
    - Ionic bond
  - The diagram below shows what type of bond?



- Ionic
- Covalent
- Metallic
- Solution

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## NAMING ACIDS

Name Key

Name the following acids.

- $HNO_3$  nitric acid
- $HCl$  hydrochloric acid
- $H_2SO_4$  sulfuric acid
- $H_2SO_3$  sulfurous acid
- $HC_2H_3O_2$  acetic acid
- $HBr$  hydrobromic acid
- $HNO_2$  nitrous acid
- $H_3PO_4$  phosphoric acid
- $H_2S$  hydrosulfuric acid
- $H_2CO_3$  carbonic acid

Write the formulas of the following acids.

- sulfuric acid  $H_2SO_4$
- nitric acid  $HNO_3$
- hydrochloric acid  $HCl$
- acetic acid  $HC_2H_3O_2$  or  $CH_3COOH$
- hydrofluoric acid  $HF$
- phosphorous acid \*  $H_2PO_3$
- carbonic acid  $H_2CO_3$
- nitrous acid  $HNO_2$
- phosphoric acid  $H_3PO_4$
- hydrosulfuric acid  $H_2S$



Name: \_\_\_\_\_ Date: \_\_\_\_\_

### Balancing Chemical Equations

Balance the following chemical equations.

- $2 \text{ Fe} + 3 \text{ H}_2\text{SO}_4 \rightarrow 1 \text{ Fe}_2(\text{SO}_4)_3 + 3 \text{ H}_2$
- $1 \text{ CH}_4 + 2 \text{ O}_2 \rightarrow 1 \text{ CO}_2 + 2 \text{ H}_2\text{O}$
- $1 \text{ SiCl}_4(l) + 2 \text{ H}_2\text{O}(l) \rightarrow 1 \text{ SiO}_2(s) + 4 \text{ HCl}(aq)$
- $2 \text{ AgI} + 1 \text{ Na}_2\text{S} \rightarrow 1 \text{ Ag}_2\text{S} + 2 \text{ NaI}$
- $4 \text{ NH}_3 + 5 \text{ O}_2 \rightarrow 4 \text{ NO} + 6 \text{ H}_2\text{O}$
- $1 \text{ FeO}_3(s) + 3 \text{ CO}(g) \rightarrow 1 \text{ Fe}(l) + 3 \text{ CO}_2(g)$
- $1 \text{ SiO}_2 + 4 \text{ HF} \rightarrow 1 \text{ SiF}_4 + 2 \text{ H}_2\text{O}$
- $2 \text{ NaBr} + 1 \text{ Cl}_2 \rightarrow 2 \text{ NaCl} + 1 \text{ Br}_2$
- $4 \text{ (NH}_4)_3\text{PO}_4 + 3 \text{ Pb(NO}_3)_4 \rightarrow 1 \text{ Pb}_3(\text{PO}_4)_4 + 12 \text{ NH}_4\text{NO}_3$
- $1 \text{ Mg(OH)}_2 + 2 \text{ HCl} \rightarrow 1 \text{ MgCl}_2 + 2 \text{ H}_2\text{O}$

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Naming ionic compounds practice worksheet name the following ionic compounds. Trying to reach a stable electron byte at least part of the time. This reaction is shown below.

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